Original Article

Electrochromic Properties of PANI/TiO\textsubscript{2} Nanocomposite Films Prepared by Electropolymerization

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Abstract: PANI/nc-TiO\textsubscript{2}/ITO composite films were prepared by electrochemical method where the monomer aniline was polymerized onto nano-porous TiO\textsubscript{2}/ITO films (PANI/nc-TiO\textsubscript{2}/ITO). The PANI/nc-TiO\textsubscript{2} heterojunctions were formed in the composite films due to the insertion of PANI in nano-porous TiO\textsubscript{2} particles. PANI/ITO films exhibited a reversible electrochromic display (ECD) performance when cycled in 0.1M LiClO\textsubscript{4} + propylene carbonate. The response times of the electrochromic coloration and bleaching of the PANI/nc-TiO\textsubscript{2}/ITO electrode were 15 s and 20 s, respectively. Electrochromic efficiency of the films reached a value as large as 12.25 cm\textsuperscript{2}\cdot C\textsuperscript{-1}. Taking advantage of the large EC efficiency and electrochemical stability as well as the simplicity of the fabrication process, PANI/nc-TiO\textsubscript{2} composite films can be used for preparation of electrochromic smart windows, in term of the large-area production in particular.

Keywords: Electrochromic device, PANI/nc-TiO\textsubscript{2}, Electrochemical deposition, Cyclic voltammogram, Coloration efficiency.

1. Introduction

Electrochromism (EC) is the phenomenon where the color or opacity of a material (usually in thin film form) changes when a polarized voltage is applied. Electrochromic smart windows, thus can block ultraviolet, visible or near infrared radiations. EC properties have been found in almost all the transition-metal oxides and their properties have been investigated extensively [1]. These oxide films can be coloured anodically (Ir, Ni) or cathodically (W, Mo). Recently Granqvist et al. [2] have made a comprehensive review of nanomaterials for benign indoor environments. In that report, the authors show
the characteristic data for a 5×5 cm² flexible EC foil incorporating WO₃, and NiO modified by the addition of a wide bandgap oxide such as MgO or Al₂O₃, PMMA-based electrolyte, and ITO films. Durability of the EC devices was demonstrated in performing several tens of thousands of coloration/bleaching cycles, and the device optical properties were found to be unchanged for many hours. Dinh et al. [3] pointed out that WO₃/TiO₂ films prepared by the doctor-blade followed by electrochemical deposition possessed both the better electrochromic performance and durability. The coloration of WO₃ deposited on indium-tin-oxides (ITO) substrates (WO₃/ITO) in 2M HCl was less than 1 sec and the maximum coloration efficiency (CE) at 630 nm was 22 cm²×C⁻¹ [4]. However, the HCl electrolyte is not suitable for practical use. The Au-doped WO₃ films were made by a dip-coating technique [5]. Beydaghyan et al [6] showed that porous and nanostructured thick WO₃ films could produce a high CE. Thummarungsan et al. [7] prepared polyaniline (PANI) copolymer films via solution casting on ITO plastic substrate. The PANI-copolymer/ITO electrode has electrochromic display of the inverse change of green-to-blue light. We also prepared PANI/ ITO by using electropolymerization of aniline in a dilute H₂SO₄ solution [8]. As-prepared PANI films have a bell-like nanoporous structure. Both the PANI-copolymer/ITO and PANI/ITO possess a coloration efficiency ranging from 8.0 to 9.0 cm²×C⁻¹. With such porous films, for a long exposed performance time, the durability of the devices was limited, making the ECD less satisfying for smart windows applications.

With the aim to improve the CE of the ECD performance and to enhance the stability of the EC devices, we used electrochemical polymerization for depositing PANI films onto nc-TiO₂. Morphology, ion exchange and electrochromic properties of the films were also presented in this paper

2. Experiment and Method

To prepare nanostructured TiO₂/ITO films, a doctor blade technique was used following the process reported in [9]. A glass slide coated with a 0.25 μm thick ITO film with a sheet resistance of 10 Ω/ and a transmittance of 90% was used as a substrate where the working area was as large as 10 cm². A colloidal solution of 15 wt. % nanoparticles (15 nm in size) of titanium oxide (Nyacol Products) in water was prepared Then the solution was filled in the slot on the ITO electrode and spread along the tapes. The samples were let for drying during 15 min, then put to a furnace maintained at 450°C for 1 hour to recrystallize the nc-TiO₂ films.

Then PANI/nc-TiO₂/ITO films were deposited by electropolymerization method using a solution of the pure reagent grade aniline and sulphuric acid (H₂SO₄) as reported in [8]. The electrochemical processes were carried-out by using a Potentiostat PGS-30 in a standard three-electrode cell, where nc-TiO₂/ITO served as working electrode (WE), Ag/AgCl as reference electrode (AAE) and a platinum grid as counter electrode (CE). The electrolyte used for the electropolymerization was composed of aqueous solution of 0.1 M aniline and 0.5M dopant sulphuric acid. PANI thin films were electropolymerized by sweeping the potential between -0.20 V and +1.20 V/AAE for 10 numbers of scans. All the experiments were performed at room temperature without stirring. The as-deposited PANI/ITO films were dried in gaseous nitrogen and kept in a clean glove-box before use. The electrochromic performance was carried out on the same PGS-30 potentiostat, using solution of 0.1M LiClO₄ in propylene carbonate for electrolyte.

The PANI films electropolymerization process in H₂SO₄ solution depends on the concentration of H₂SO₄, the scan rate and the range of the potentials used. Figure 1 shows typical cyclic voltammograms (CVs) of the best electropolymerized PANI films. The electrolyte contains 0.1M aniline monomer with 0.5M concentration of H₂SO₄. The CVs were recorded for 10 successive cycles at a scan rate of 50 mV/s for potential range from -0.1 to +0.6 V/AAE. We observed the decrease in growth rate upon repeating
the potential cycle [10] where the authors explained this phenomenon by the poor electrochemical activity of the PANI film. After completing 10 cycles, greenish-colored thin films were deposited which confirms the formation of emeraldine state (ES) of PANI.

3. Result and Discussion

3.1. Film Morphology

Bright-field micrographs of the samples is shown in Figure 2. The thickness of the composite film was determined by the point-to-point technique, \(d \approx 450\) nm. The surface of the doctor-blade deposited sample clearly shows the nanoscale porosity of the TiO\(_2\) film (Figure 2a). The electropolymerization of PANI resulted in the filling-up of PANI chains into the pores of the nanostructured TiO\(_2\) structure. As a result, numerous PANI/nc-TiO\(_2\) heterojunctions were created in the composite film (Figure 2b).
3.2. Electrochemical Property

Figure 3 shows the cyclic voltammetry (CV) curve in 0.1MLiClO$_4$ + PC of a PANI/nc-TiO$_2$/ITO film served as the working electrode (WE) where the CV spectra was recorded at the fifth cycle. This is a typical curve for the films of 450nm thickness prepared by our studies.

From this figure one can see a nearly symmetrical shape of the CV spectra toward the coordinates ($x = 0.60; y = 0.16$). In the positive sweep direction (PSD) a peak of the anodic current density corresponding to a value of ca. 0.55 mA was obtained at a potential of 1.05 V/AAE. A negative value ($\sim 0.21$ mA) of the peak in the negative sweep direction (NSD) was obtained at a potential of 0.36 V/AAE. The symmetrical CV proves a good reversibility of the processes of (ClO$_4^-$) ion insertion/extraction from the electrolyte into/out of the working electrode (WE). Since PANI/nc-TiO$_2$/ITO consists of two electrochromic materials, namely polymeric PANI and inorganic nc-TiO$_2$, electrochromic performance can be expressed by a double overall reaction, as follows.

![Figure 3. Cyclic voltammetry spectra of PANI/nc-TiO$_2$/ITO cycled in LiClO$_4$+PC (scanning rate $\mu$ = 100 mV/s).](image)

During the oxidation (PANI/nc-TiO$_2$/ITO is positively polarized), the ClO$_4^-$ ions from LiClO$_4$/PC electrolyte solution insert into and Li$^+$ ions extract out of the WE. The first double reaction (corresponding to the bleaching state) can be expressed as [11]:

$$\text{(ANI)}_n + ny(\text{ClO}_4^-) \rightarrow [(\text{ANI}^+) + y(\text{ClO}_4^-)]_n + ny_e^- \quad (1)$$

where $n$ is the number of repeated units and $y$ is the stoichiometric number of the counter ion; and as [12,13]:

$$\text{Li}_x\text{TiO}_2 \rightarrow \text{TiO}_2 + x(\text{Li}^+ + e^-) \quad (2)$$

During the reduction (PANI/nc-TiO$_2$/ITO is negatively polarized), the ClO$_4^-$ ions of LiClO$_4$/PC electrolyte solution were excluded from the WE, whereas Li$^+$ ions were inserted into it. The second double reaction (corresponding to the coloration state) is described by two above equations with the inverse.
direction of the arrows. Thus, combining Eqs. (1) and (2), one can describe the electrochromic performance of PANI/nc-TiO$_2$/ITO electrode by an overall reaction, as follows:

$$\text{PANI-CIO}_4^-/\text{TiO}_2 + \text{Li}^+ \leftrightarrow \text{PANI/Li}_x\text{TiO}_2 + \text{ClO}_4^-$$

(Light green) (Blue)

The electrochromic coloration of the PANI/nc-TiO$_2$/ITO occurred due to the simultaneous double possesses of insertion and extraction of ClO$_4^-$ and Li$^+$ ions, respectively. In our case, for the composite WE during the performance, the electrochromic display is due to a reversible light-green and blue colour change. This is a specific difference in PANI/nc-TiO$_2$ based EC devices compared to that of the single WO$_3$ or TiO$_2$ films displaying a reversible transparent-to-blue change.

3.3. Electrochromic Performance

For a sample with a 450 nm-thick PANI/nc-TiO$_2$ film (WE), the in situ transmission spectra, obtained during coloration at a polarized potential of −0.5 V/AAE are given in Figure 4.

![Figure 4](image)

Figure 4. Transmittances of the PANI/nc-TiO$_2$ film. The oxidizing potential and reducing potential were set +0.3 V/AAE and -0.5 V/AAE, respectively. Curve “a” is the colored state and curve “b” is bleached state.

From Figure 4 one can see a large difference in transmittance spectra in the visible range of the film between the colored (curve “a”) and bleached (curve “b”) states. In about 15s, the transmittance at $\lambda = 550$ nm from ~ 70% decreased to ~ 16% with the colored state, corresponding to bias potential − 0.5 V/AAE. With a polarized potential of + 0.3 V/AAE, the WE bleached in 20s. This result indicates that the electrochemically deposited PANI/nc-TiO$_2$ films exhibited a good reversibility of the EC performance. From the above mentioned results for the ECDs with heterojunctions of PANI/nc-TiO$_2$, it is seen that the efficient coloration can be achieved due to a double coloration process, as shown in Eqs. (1) and (2).

The electrochromic coloration efficiency ($\eta$) was calculated using well-known expression relating the efficiency with the optical density, consequently the transmittances of coloration ($T_c$) and bleaching states ($T_b$), and the insertion charge ($Q$), as follows [14]:

$$\eta = \frac{\Delta OD}{Q} = \frac{1}{Q} \ln \left( \frac{T_b}{T_c} \right)$$

(4)
In our experiments $Q=0.12 \text{ mC} \times \text{cm}^2$. At a wavelength of 550 nm, $T_s=70\%$ and $T_c=16\%$, thus the coloration efficiency was found to be of $\sim 12.25 \text{ cm}^2 \times \text{C}^{-1}$. In the visible range of wavelengths, $\eta$ of the PANI/nc-TiO$_2$ device was found to be larger than that of the PANI/ITO device made by electropolymerization [8] as well as made by electrochemistry [15].

4. Conclusion

PANI/nc-TiO$_2$/ITO composite films were prepared by electropolymerization of monomer aniline onto nc-TiO$_2$/ITO films which were made from the doctor-blade technique. By embedding PANI in porous TiO$_2$ nanoparticles, PANI/nc-TiO$_2$ heterojunctions were formed in the composite films. The coloration efficiency and electrochemical stability of the PANI/nc-TiO$_2$ films were considerably improved. The response times of the ECD coloration/bleaching and electrochromic efficiency of PANI/nc-TiO$_2$/ITO were 15–20 s and 12.25 cm$^2$/C$^{-1}$, respectively. Since a large-area PANI/nc-TiO$_2$/ITO electrode can be prepared by the doctor-blade method, the PANI/nc-TiO$_2$/ITO films contribute a good candidate for smart window applications, taking advantage of its large efficiency and electrochemical stability as well as the simplicity of the fabrication process.

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